**Ionization Energy, and Electron Affinity, and Electronegativity**

1. Explain why the ionization energy of lithium is more than beryllium.
2. Explain why fluorine is more electronegative than bromine.
3. Argon has a lower electron affinity than chlorine. Discuss why this must be true.
4. Distinguish between the 1st, 2nd, and 3rd ionization energies of an atom.
5. Explain why the ionization energy of magnesium is higher than strontium.
6. Which atom in each of the following pairs would have the lower first ionization energy?

a. N, O f. Mg, Ca

b. Te, Sn g. I, Sb

c. Ne, F h. Al, N

d. C, Ge i. F, S

e. Br, I

1. Which of the following has a higher electron affinity?

a. Br, Ca e. V, Ta

b. O, Te f. Na, P

c. Fr, Rb g. F, At

d. Al, Tl h. Zr, Sb

1. Which of the following has a higher electronegativity?

a. Fr, K

b. F, B

c. Te, S

d. Sr, Be