1. What type of intermolecular forces would you expect to find in a sample of PF3?

a. London

b. covalent

c. interionic

d. H-bonding

e. dipole-dipole

2. What are the strongest intermolecular forces present in a liquid sample of NH3?

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

3. What are the strongest intermolecular forces present in a liquid sample of PCl3?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

4. What type of intermolecular forces would you expect to be the predominant (strongest) forces in a liquid sample of H2S?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

5. What type of intermolecular forces wouldyou expect to be the predominant (strongest)forces in a liquid sample of CBr2F2?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

6. What accounts for the strongest attractions between molecules of H2Se?

a. the partial positive charge on the H from one molecule is attracted to the partial negative charge on the Se from another molecule.

b. The −1 charge of the hydrogen ion is attracted to the +2 charge on the selenium ion.

c. The +1 charge of the hydrogen ion is attracted to the −2 charge on the selenium ion.

d. the partial negative charge on the H from one molecule is attracted to the partial positive charge on the Se from another molecule.

7. What type of intermolecular forces would you expect to be the predominant (strongest) forces in a sample of Xe?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

8. What type of intermolecular forces would you expect to be the predominant (strongest) forces in a sample of SrF2?

a. dipole-dipole

b. hydrogen bonding

c. dispersion

d. covalent

e. ion-ion

What type of intermolecular forces would you expect to find in a liquid sample of HBr?\_\_\_\_\_\_\_\_\_\_\_\_

Which type of intermolecular forces or intramolecular forces are the strongest forces contributing to the melting point of pure HF?

a. covalent

b. ionic

c. dispersion

d. metallic

e. dipole-dipole

f. hydrogen-bonding

Key

1. E

2. H-bonding

3. dipole-dipole

4. dipole-dipole

5. dipole-diponle

6. A

7. London dispersion

8. E

9. dipole-dipole

10. H-bonding