**Check sheet for AP**

* HW problems 13 and 14 (in notebook)
* Study sheet (in notebook)
* Pre-lab
* Data table
* Post lab questions
* Appropriate HW #s (according to graded solutions quiz)
* Organize notebook (page #s and table of contents)

**Stoichiometry Practice**

For each problem you MUST

1. Write the correct and complete balanced equation
2. Show your work including units on EVERY SINGLE number.
3. How many grams of aluminum are required to react with 35 mL of 2.0 *M* hydrochloric acid, HCl?
4. How many grams of sodium can be reacted with 750 mL of a 6.0 *M* solution of sulfuric acid, H2SO4?
5. If 45 mL of a 1.5 *M* AgNO3 is added to 1.5 grams of KCl how many grams of AgCl can be formed?
6. How many liters of a 0.75 *M* solution of Ca(NO3)2 will be required to react with 148 g of Na2CO3?
7. How many liters of a 3.0 *M* H3PO4 solution are required to react with 4.5 g of zinc?
8. How many milliliters of 0.10 *M* Pb(NO3)2 are required to react with 75 mL of 0.20 *M* NaI?
9. How many grams of solid barium sulfate will form when 25 mL of 0.33 M sodium sulfate reacts with 25 mL of 0.50 *M* barium nitrate?
10. If 525 mL of 0.80 *M* HCl solution is neutralized with 315 mL of Sr(OH)2 solution what is the molarity of the Sr(OH)2?

**REDOX practice**

Before you begin, write the steps for balancing REDOX reactions at the top of the paper (all of them).

As you work the problems, check off that you have completed each step (may want to use a different color check for each problem).

1. Mn 2+ + BiO3 - ===> MnO4 - + Bi 3+
2. MnO4 - + S2O3 2- ===> S4O6 2- + Mn 2+
3. P + Cu 2+ ===> Cu + H2PO4 –
4. PH3 + I2 ===> H3PO2 - + I -