Redox Titration Post-lab

1. Complete **Calculations** as described on page 57.

2. If you were able to do more than one trial, calculate the average molarity from the trials.

3. What was the purpose of the addition of the acid to the solution?

4. In a titration, the **analyte** is the substance being examined (unknown molarity), while the **titrant** is the substance added to the analyte that has a known concentration. Write these definitions in your own words, then identify the analyte and titrant in this experiment.

5. Several errors could have been made in the lab. Discuss how these would have affected the final calculated concentration of your KMnO4 solution (higher that it should have been, lower than it should have been, no change).

a. The burette was not rinsed with the solution prior to filling. Some water was left in the burette when it was filled.

b. KMnO4 was added to the Erlenmeyer flask until a deep purpose or red color appeared.

c. More than 25 mL of water was added to the FAS when the solution was created.

d. Some FAS was spilled as it was transferred from the scale to the container.

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